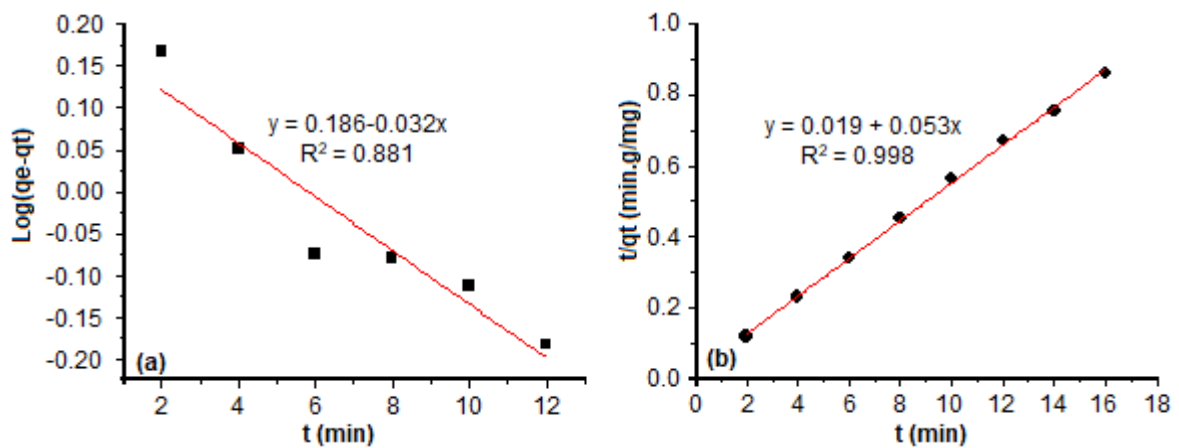


### Supplementary Data

This supplementary data is a part of paper entitled “Adsorption of Remazol Brilliant Blue R Using Amino-Functionalized Organosilane in Aqueous Solution”.

#### Adsorption Kinetic

From the adsorption data of RBBR onto AFOS was used to calculate the kinetic parameters such as pseudo first-order and pseudo second-order to obtain information of  $k_1$ ,  $k_2$  and which model is fitted to the adsorption. Fig. S1(a) shows the pseudo first-order kinetic plot linier while Fig. S1(b) present the pseudo second-order plot linier. From the  $R^2$  values, the pseudo second-order is fitted to the adsorption.

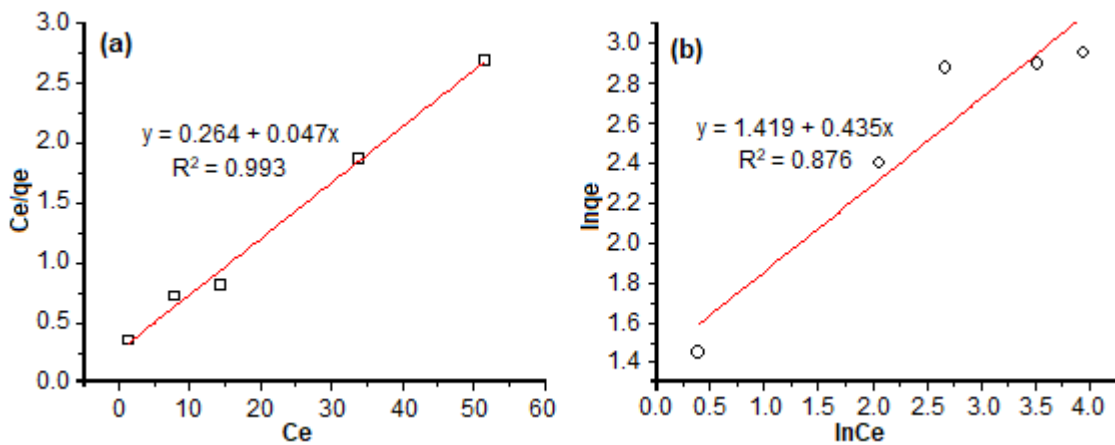


**Fig S1.** (a) Pseudo first-order and (b) Pseudo second-order kinetics for the adsorption of RBBR onto AFOS

**Isotherm Adsorption**

From Langmuir Equation below, it can be fitted Ce versus Ce/Qe plot shown in Fig. S2(a)

$$\frac{Ce}{Qe} = \frac{1}{K_L \times Q_m} + \frac{1}{Q_m} Ce$$



**Fig S2.** Isotherm (a) Langmuir and (b) Freundlich models plot curve

Based on Ce versus Ce/Qe graph can be obtained linier equation  $y=0.047x + 0.264$ . From this data, the Qm (monolayer adsorption capacity, mg/g) and Langmuir constant can be calculated as presented in Table 2 in main paper.

From Freundlich equation showed below, can be plotted lnCe versus lnQe graph shown in Fig. S2(b).

$$\ln Qe = \ln K_f + \frac{1}{n} \ln Ce$$

It was obtained a linier equation from lnCe versus lnQe plot graph,  $y=0.435x + 1.419$ . This data was used to measure the n and kf values shown in Table 2 in main paper.