

Styrene and Azo-Styrene Based Colorimetric Sensors for Highly Selective Detection of Cyanide

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ABSTRACT

A novel styrene (1) and azo-styrene (2) based chemosensor from vanillin has been successfully synthesized. Sensor 1 was obtained via one step Knoevenagel condensation using Ultrasound method and sensor 2 was afforded via azo coupling and Knoevenagel condensation reactions. Both of sensors showed high sensitivity and selectivity to detect CN^- in aqueous media, even in the presence of other anions, such as F^- , Cl^- , Br^- , I^- , CO_3^{2-} , SO_4^{2-} , H_2PO_4^- , and AcO^- . Colorimetric sensing of sensor 1 was due to the deprotonation of sensor binding site (-OH) by cyanide anion. It consequently improved the selectivity and sensitivity in sensing cyanide which was observed by the color change of sensor from yellowness to red-purple and then purple in 1 min. Sensor 2 exhibited chemodosimeter phenomenon which is caused by the nucleophilic addition of cyanide and gave color change from green to greenish-blue. The results enable to detect an detection by naked eye.

Keywords: colorimetric chemosensor; styrene; azo-styrene; cyanide

ABSTRAK

Sintesis senyawa turunan stirena (1) dan azo-stirena (2) dari vanilin telah berhasil dikembangkan sebagai kemosensor anion. Sensor 1 diperoleh dari reaksi kondensasi Knoevenagel, sedangkan sensor 2 diperoleh dari reaksi koping diazo dan kondensasi Knoevenagel. Uji aktivitas senyawa 1 dan 2 dilakukan pada anion F^- , Cl^- , Br^- , I^- , CO_3^{2-} , SO_4^{2-} , H_2PO_4^- , CN^- , and AcO^- . Hasil penelitian menunjukkan bahwa sensor 1 dan 2 memiliki sensitivitas dan selektivitas terhadap anion CN^- yang memberikan perubahan warna dari kekuningan menjadi merah-ungu kemudian ungu dalam waktu 1 menit, sementara sensor 2 menghasilkan perubahan warna dari hijau menjadi biru-kehijauan. Perubahan warna pada sensor 1 disebabkan oleh deprotonasi anion CN^- pada sisi ikat -OH, sedangkan pada sensor 2 perubahan warna terjadi dikarenakan adisi nukleofilik oleh anion sianida yang ditunjukkan dari spektra UV-vis mengalami pergeseran hipsokromik. Proses deteksi anion pada sensor 1 dan 2 menghasilkan perubahan warna yang dapat diamati secara kasat mata, sehingga mempermudah deteksi anion.

Kata Kunci: kolorimetri; stirena; azo-stirena; anion sianida

INTRODUCTION

The recognition and sensing of anions are the growing interest area in supramolecular chemistry due to its important role in a wide range of chemical, biological, environmental, and industrial applications [1]. Among the various anions (i.e. fluoride, chloride, acetate, and phosphate), cyanide (CN^-) ion gained great interest because of its lethal effect on living organisms and environment [2]. Nevertheless, CN^- is widespread in an industrial process, such as gold mining, electroplating, metallurgy, and the synthesis of nylon, fibers, and resins which have resulted in severe pollution of water supplies [3]. Cyanides also do harm to human beings by absorption through the lungs, gastrointestinal tract and skin, leading to vomiting, convulsions, loss of consciousness, and eventual death [4]. Therefore, highly

selective and sensitive probes for the cyanide anion are of considerable interest due to their application to environmental and the pathological imaging of the anions [5].

Many methods have been used for anion recognition, but most of them are not simple, relatively expensive, and need a long time in detection. Naked eye detection of anions has attracted much attention and offers the advantage such as quick response, monitoring output visually, no need of any sophisticated equipment and has high sensitivity [6]. Generally, sensor compound consists of the binding site and signalling unit (chromogenic and fluorogenic). The signalling unit is attached covalently or non-covalently with the binding site [7]. Neutral sensors have been used for anions detection and most of them contain

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amide, indocarbazole, imidazolium, urea, thiourea, hydrazone, pyrrole [8].

Currently, many excellent anion chemosensors have been reported with high selectivity and sensitivity. However, there are disadvantageous due to the complicated structure of those chemosensors, requiring of hard synthetic routes or troubling in the purification process and give poor yield. Shao et al. have reported naked eye and fluorogenic detection of cyanide using urea-coumarin derivatives based on intramolecular charge transfer (ICT) which is not easy to synthesize and merely showed selective response only in the presence of acetate ion [8]. Velmathy et al. also have reported fluorogenic and chromogenic heterocyclic pyrrole derivatives to recognize CN⁻ ion which is also showed complicated compound to synthesize [9].

Colorimetric anion sensor design, with high sensitivity and selectivity that both easy to synthesize and to recognize by the naked eye, become a new challenge. In this research, we have developed styrene **1** and azo-styrene **2** derivatives from vanillin via Knoevenagel condensation that were synthesized using ultrasound method [10]. Vanillin was used as starting material because it has a hydroxyl group as a binding site and aldehyde group which can be modified as a signalling unit. In this project, sensor **1** contained styrene group that bind nitro group and sensor **2** was synthesized as the modification of sensor **1**. In addition, the nitro group as chromophore unit can increase the ability of the sensor to bind with anions. Azo group was added to sensor **2** as chromophore unit because it can extend delocalization and increase the ability to recognize anion by the naked eye. This result will give high selectivity of anion sensors.

EXPERIMENTAL SECTION

Materials

Vanillin, nitromethane, glacial acetic acid, ammonium acetate, *p*-nitroaniline, hydrochloric acid, aquadest, sodium azide, sodium nitrite, ethanol, methanol, dimethyl sulfoxide (DMSO), acetone, sodium fluoride, sodium bromide, sodium chloride, sodium iodide, sodium carbonate, sodium sulphate, sodium dihydrogen phosphate, sodium cyanide, and sodium acetate. All reagents in analytical grade were obtained from E Merck Co Inc. (Germany) and were used without further purification.

Instrumentation

In general, the melting point of the compounds was determined using melting point electrothermal 9100.. Synthesis was performed by using Ultrasound Bruder

1200. FTIR spectra were recorded using Shimadzu Prestige-21 instrument use KBr pellet. GC-MS were recorded using Shimadzu QP2010S. The ¹H-NMR and ¹³C-NMR spectra were recorded from Nuclear Magnetic Resonance Spectrometer (JEOL-MY500). UV-vis spectra were recorded with Shimadzu UV-1800 Eng 240 V spectrophotometer with a quartz cuvette (path length = 1 cm) at room temperature (r.t). For all spectroscopic titration, 5×10⁻⁵ M concentrations were prepared in DMSO and saturated anions were prepared in aqueous medium.

Procedure

Synthesis and characterization of sensor 1

Sensor **1** was synthesized via ultrasound methods. Into mixture of nitromethane (6.50 mL) and ammonium acetate (1.65 g) was added to a mixture of glacial acetic acid (1.65 mL) and vanillin (1.52 g, 10 mmol) then sonicated at room temperature for 2 h. The precipitate was filtered followed by washing with aquadest then dried. The product was purified by recrystallization from methanol Sensor **1** was obtained as a orange solid in 70.87%, m.p. 166.2 °C; FT-IR (KBr) (cm⁻¹): 3471 (-OH group), 1604 (C=C alkyne), 1519 (C=C aromatic), 1489 and 1288 (-NO₂ group), 1126 and 1018 (-C-O-C), 954 (-C=C trans), m/z [M⁺]: 195.

Synthesis and characterization of sensor 2

Sensor **2** was synthesized via coupling diazo and Knoevenagel condensation reactions by ultrasound method.

Azo coupling. Into solution of hydrochloric acid 37% (7.40 mL, 75 mmol) and aquadest (7.40 mL) was added *p*-nitroaniline (3.45 g, 25 mmol). The mixture was stirred in 0-5 °C. To the stirred solution was added solution of sodium nitrite (2.10 g, 30 mmol) in 10 mL aquadest. The mixture was stirred for 1 h in 0-5 °C. The product was solution of benzene diazonium chloride. Into a solution of sodium carbonate (10 g) and aquadest (75 mL) was added vanillin (3.80 g, 25 mmol). The mixture was cooled at 0-5 °C. Into benzene diazonium chloride solution was added vanillin solution drops for 1 h and stirred the mixture for 4 h. The precipitate was neutralized by hydrochloride acid solution 10% then was filtered followed by washing with aquadest then was dried to give the desired compound as brown solid. The crude was purified by recrystallization with ethanol. FTIR (KBr) v (cm⁻¹) 3448 (OH group), 3109 (Csp²-H), 1689 (C=O), 1604 (C=C aromatic), 1519 and 1342 (-NO₂), 1427 (-N=N), 1280 and 1137 (C-O-C).

Knoevenagel condensation reaction. Into mixture of nitromethane (6.50 mL), ammonium acetate (1.65 g),

and glacial acetic acid (1.70 mL) was added azo-vanillin (0.37 g, 1.25 mmol). The reaction mixture was sonicated at room temperature for 3 h. The precipitate was filtered followed by washing with aquadest then dried. The product was purified by recrystallization from methanol as brown powder in 81.40%, m.p. 218; FTIR (KBr) ν (cm^{-1}) 3488 (OH group), 3116 ($\text{Csp}^2\text{-H}$), 1620 (C=C alkyne), 1604 (C=C aromatic), 1489 and 1342 ($-\text{NO}_2$), 1427 ($-\text{N=N}$), 1157 and 1033 (C-O-C), 972 (C=C trans); $^1\text{H-NMR}$ (500 MHz, DMSO-d_6) δ (ppm): 3.94 (1H, s, OCH_3), 7.67 (1H, s, ArH_1), 7.77 (1H, s, ArH_1), 8.12-8.15 (2H, d of d, ArH_2), 8.20-8.22 (1H, d, $-\text{C=C}$), 8.28-8.30 (2H, d of d, ArH_2), 8.43-8.45 (1H, d, $-\text{C=C}$); $^{13}\text{C-NMR}$ δ (ppm): 56.49 ($-\text{OCH}_3$), 113.91 ($\text{Ar}_1\text{-CH}$), 115.26 ($\text{Ar}_1\text{-CH}$), 121.26 ($\text{Ar}_2\text{-CH}$), 123.84 ($\text{Ar}_2\text{-CH}$), 125.04 ($\text{Ar}_1\text{-C}$), 136.88 ($\text{Ar}_1\text{-C}$), 139.19 ($\text{Ar}_1\text{-C}$), 139.53 (C=C), 148.34 (C=C), 150.19 ($\text{Ar}_2\text{-C-NO}_2$), 151.07 ($\text{Ar}_1\text{-C-OMe}$), 154.93 ($\text{Ar}_2\text{-C-N=N}$).

Study activities solvatochromic and ionochromic of sensor 1 and 2

Solvatochromic experiment. As much as 2 mL solution of the sensor 1 and 2 in DMSO were prepared and the colour change was observed, then UV-vis spectra were measured in concentration of 5.0×10^{-5} M.

Ionochromic experiment. All titration experiments were carried out at room temperature, unless otherwise mentioned. UV-vis spectra was measured using a Shimadzu UV-1800 Eng 240 V spectrophotometer. As much as of 2 mL, 5.0×10^{-5} M solution of the sensor 1 and 2 in DMSO and solutions of saturated salt anions

were prepared in aquadest. Into 2 mL of sensor 1 and 2 was added three drops of anions. These solutions were used for all spectroscopic studies after appropriate dilution.

LoD calculation

Sensor 1. Based on Table 2 standar deviation value = 0.00427 and Fig.13, slope value (m) = 1.375, so limit of detection (LOD) sensor 1:

$$\begin{aligned} \text{LOD} &= 3\sigma/m \\ &= 3(0.00427)/1.375 \\ &= 9.316 \times 10^{-3} \text{ M} \end{aligned}$$

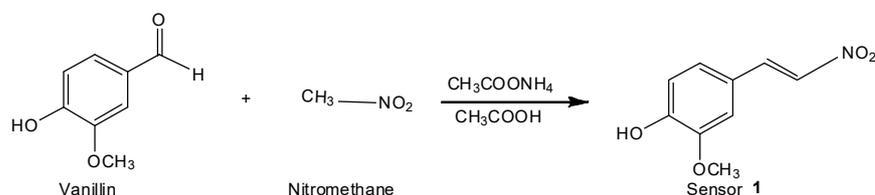
Sensor 2. Based on Table 4 standar deviation value = 0.10215 and Fig.16, slope value (m) = 22.199, so limit of detection (LOD) sensor 2:

$$\begin{aligned} \text{LOD} &= 3\sigma/m \\ &= 3(0.10215)/22.199 \\ &= 1.370 \times 10^{-2} \text{ M} \end{aligned}$$

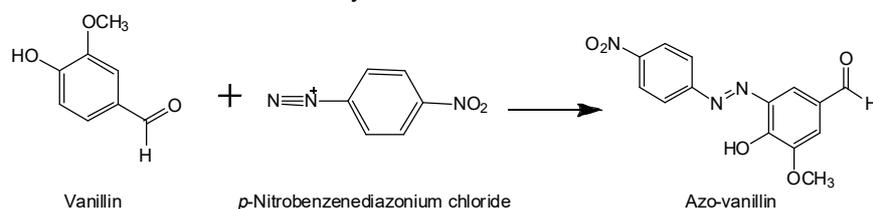
RESULT AND DISCUSSION

Synthesis of Sensor 1

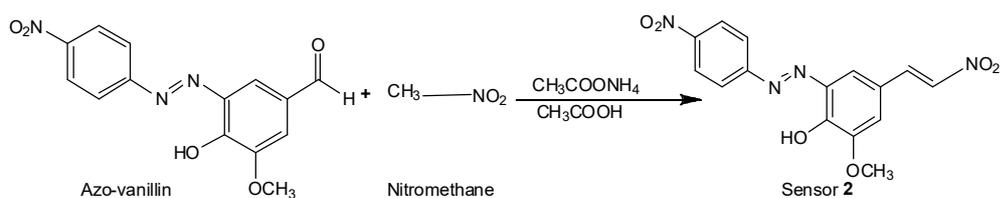
Vanillin contains aldehyde group which can be modified as styrene group via Knoevenagel condensation that binds nitro group. This reaction was occurred in the presence of ammonium acetate and glacial acetic acid as catalyst without solvent (Scheme 1). Ammonium acetate will deprotonate nitromethane



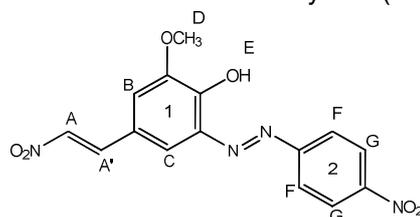
Scheme 1. Synthesis of sensor 1 from vanillin



Scheme 2. Synthesis of azo-vanillin via coupling diazo



Scheme 3. Synthesis of sensor 2

Table 1. ¹H-NMR data of azo-styrene (sensor 2)

Signal	δ (ppm)	Multiplicity	Integration	Proton (Hx)
1	3.940	Singlet	3H	D
2	7.674	Singlet	1H	C
3	7.777	Singlet	1H	B
4	8.129-8.155	Doublet ($J = 13$ Hz)	2H	F
5	8.205-8.452	Doublet	1H	A'
6	8.282-8.308	Doublet ($J = 13$ Hz)	2H	G
7	8.436-8.452	Doublet ($J = 8$ Hz)	1H	A

Table 2. Blank data of sensor 1

No	Sample	Concentration (M)	Absorbance 540 nm
1	Blank 1	5×10^{-5}	0.009
2	Blank 2	5×10^{-5}	0.004
3	Blank 3	5×10^{-5}	0.003
4	Blank 4	5×10^{-5}	0.013
5	Blank 5	5×10^{-5}	0.013
Standard deviation			0.00427

Table 3. LOD data of sensor 1

No	Sensor 1 + CN ⁻ (equivalent)	Absorbance		A540/A391
		391 (nm)	540 (nm)	
1	10^{-5}	0.678	0.114	0.168
2	10^{-4}	0.668	0.107	0.186
3	10^{-3}	0.653	0.107	0.190
4	10^{-2}	0.626	0.200	0.319
5	10^{-1}	0.352	1.042	2.960
6	1	0.085	1.839	21.635
7	5	0.066	1.807	27.387
8	10	0.039	1.863	47.769
9	30	0.023	1.805	78.478
10	50	0.017	1.819	97.470
11	70	0.017	1.657	107.000
12	90	0.014	1.640	117.142

become nucleophile because nitromethane is stable compound, so it needs to be changed become more active. Glacial acetic acid plays as a catalyst in this reaction.

The reaction process was observed and monitored by TLC with eluent ratio of ethyl acetate:*n*-hexane = 1:3. The product of this reaction was ((*E*)-2-methoxy-4-(2-nitrovinyl)phenol (1) that obtained as orange powder in 70.87% yield and melting point of 166.2 °C. Structural determination of the product was based on FTIR and GC-MS studies. The FTIR spectrum showed several peaks of -OH, -OCH₃, -NO₂, C-H aromatic, C=C aromatic, C=C alkyne, C_{sp2}-H aliphatic, and C_{sp3}-aliphatic groups. The most important information from the FTIR spectrum was the absence of aldehyde group signal at

1666 cm⁻¹ for the starting material of vanillin. Besides, the presence of styrene group signal at 1604 cm⁻¹ and nitro group signals with sharp twin peaks at 1489 and 1288 cm⁻¹.

Synthesis of Sensor 2

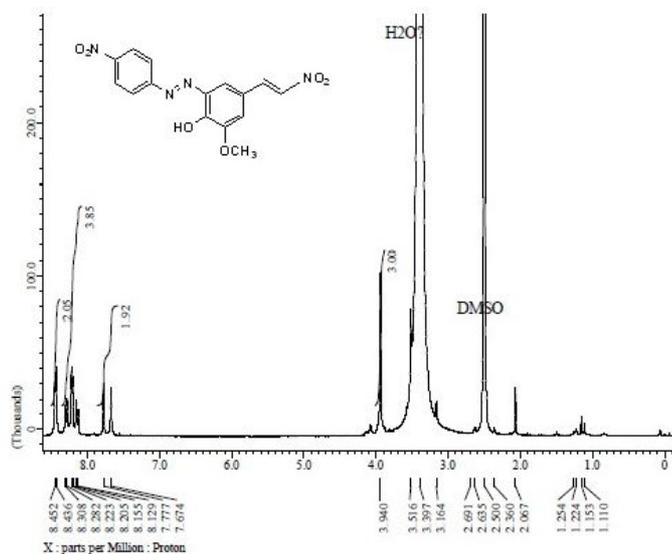
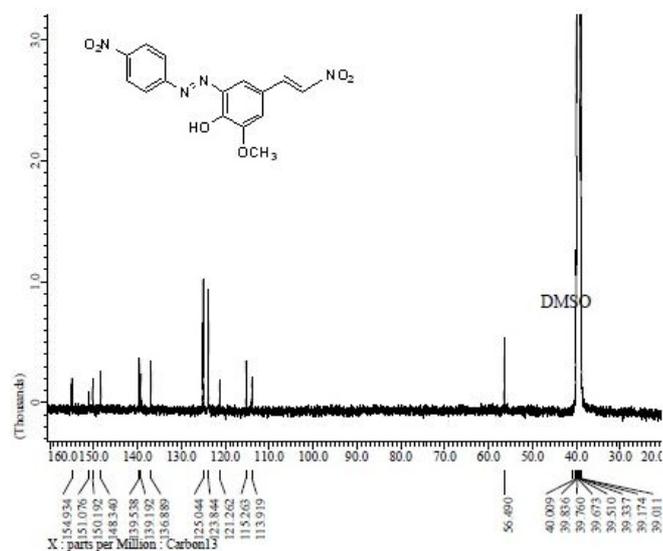
Sensor 2 can be obtained by two-step reaction, that is azo coupling and Knoevenagel condensation. Azo coupling reaction occurred between vanillin and *p*-nitroaniline to product azo-vanillin (Scheme 2). This reaction was began with formation of vanillat ion from vanillin in the presence of sodium hydroxide, so electron density of vanillin ring was increased. This condition caused benzene ring became more active,

Table 4. Blank data of sensor 2

No	Sample	Concentration (M)	Absorbance 633 (nm)
1	Blank 1	5×10^{-5}	0.600
2	Blank 2	5×10^{-5}	0.341
3	Blank 3	5×10^{-5}	0.538
4	Blank 4	5×10^{-5}	0.371
5	Blank 5	5×10^{-5}	0.388
Standard deviation			0.10215

Table 5. LOD data of sensor 2

No	Sensor 1 + CN ⁻ (equivalent)	Absorbance		A633/A385
		385 (nm)	633 (nm)	
1	10^{-5}	1.146	0.655	0.571
2	10^{-4}	0.975	0.674	0.691
3	10^{-3}	1.019	0.599	0.587
4	10^{-2}	0.881	0.852	0.967
5	10^{-1}	0.500	1.425	2.850

**Fig 1.** ¹H-NMR spectrum of azo-styrene (2)**Fig 2.** ¹³C-NMR spectrum of azo-styrene (2)

therefore, it can easily react with benzenediazonium ion.

This reaction process was observed and monitored by TLC with eluent ratio of ethyl acetate:*n*-hexane = 2:3. The product of this reaction was ((*E*)-1-(4-hydroxy-3-methoxy-5-((4-nitrophenyl)diazenyl)-phenyl)ethan-1-on(2) that was obtained as brown powder in 68.22% yield with melting point 231.2-232.8 °C. Structural determination of the product was made on the basis of FTIR and GC-MS studies. The FTIR spectrum of this compound showed many peaks consists of -OH, -NO₂, azo (N=N), -OCH₃, C=O, C-H aromatic and C=C aromatic groups. The most important information of FTIR spectrum is the presence of azo group signal at 1427 cm⁻¹ which indicated the product of azo-vanillin was generated.

Azo-styrene compound (sensor 2) can be obtained via Knoevenagel condensation reaction between azo-

vanillin and nitromethane in presence of ammonium acetate and glacial acetic acid. This reaction was easier because it was carry out at room temperature via ultrasound method. Mechanisms of this reaction are shown in Scheme 3. This reaction process was observed and monitored by TLC with eluent ratio of ethyl acetate:*n*-hexane = 2:3. The product of this reaction was brown powder in 81.40% yield with melting point at 218-219 °C. The FTIR spectrum showed many peaks consist of styrene (C=C alkyne), -OH, -NO₂, azo (N=N), -OCH₃, C-H aromatic and C=C aromatic groups. The most important information of this spectrum was the absence of aldehyde group signal at 1689 cm⁻¹ for starting material of azo-vanillin. Besides, the presence of styrene (C=C alkyne) group signal at 1620 cm⁻¹ that indicated the product was formed.

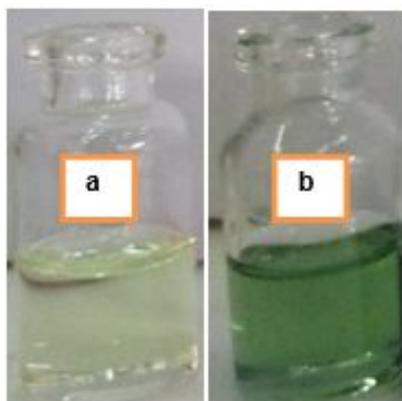


Fig 3. Solvatochromic phenomena target compound ($5 \times 10^{-3} \text{M}$) in DMSO: a. Sensor 1, b. Sensor 2

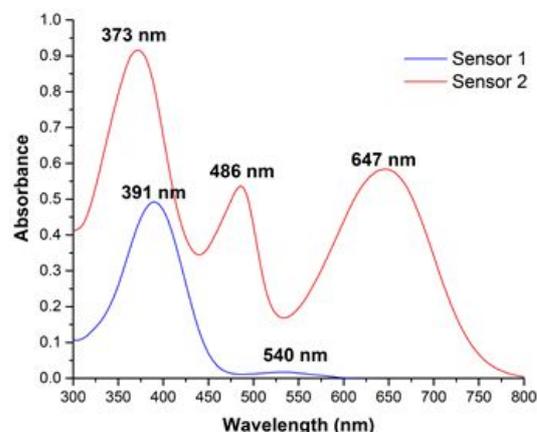


Fig 4. Spectra absorption of sensor 1 and 2 in DMSO

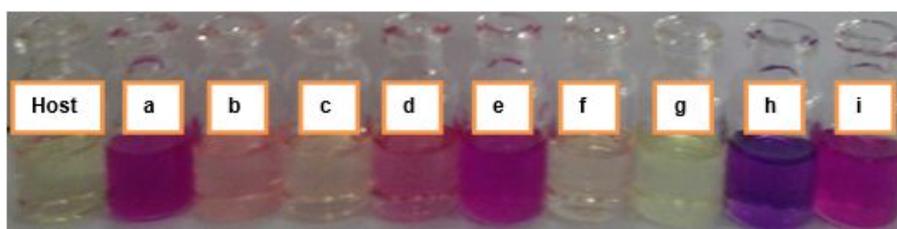


Fig 5. Color change of sensor 1 in DMSO ($5 \times 10^{-5} \text{M}$): a. F^- , b. Cl^- , c. Br^- , d. I^- , e. CO_3^{2-} , f. SO_4^{2-} , g. H_2PO_4^- , h. CN^- , i. AcO^-



Fig 6. Color change of sensor 2 in DMSO ($5 \times 10^{-5} \text{M}$): a. F^- , b. Cl^- , c. Br^- , d. I^- , e. CO_3^{2-} , f. SO_4^{2-} , g. H_2PO_4^- , h. CN^- , i. AcO^-

In the $^1\text{H-NMR}$ spectrum (Fig. 1) displayed seven peaks that is interpreted in Table 1. The clear evidence for the azo-styrene product was given by $^{13}\text{C-NMR}$ spectrum (Fig. 2) in which typical signal corresponding to styrene carbon ($-\text{C}=\text{C}-\text{NO}_2$) and ($\text{Ar}_1-\text{C}=\text{C}-$) appeared at 139.53 ppm and 148.34 ppm, respectively. MS (m/z): 344 [M^+].

Study Activities Solvatochromic and Ionochromic of Sensor 1 and 2

Solvatochromic experiment

Solvatochromic phenomena of sensor 1 and 2 in DMSO are shown in Fig. 3 and absorption spectra are shown in Fig. 4. Based on Fig. 3 and 4, there is a relationship between color change of sensors and UV-vis absorption spectra. DMSO gives sharper color in sensor 2 than sensor 1. Color of sensor 1 was yellowish,

whereas that of sensor 2 was green. This result was confirmed by UV-vis spectrophotometer [11-12].

Based on Fig. 4, there are two kind electron transition of sensor 1 and 2 that is $n \rightarrow \pi^*$ and $\pi \rightarrow \pi^*$. Sensor 2 has a maximum wavelength at 373, 486 and 647 nm meanwhile sensor 1 only has 391 and 540 nm. Sensor 2 showed more significant bathochromic shift than sensor 1 because it has long electron delocalization probably caused by presence of azo group and two nitro group as chromogenic unit, so sensor 2 shows sharp color than sensor 1.

Ionochromic experiments

The sensing properties of sensor 1 and 2 were examined in saturated salt anion in H_2O . Some examined anions were F^- , Cl^- , Br^- , I^- , CO_3^{2-} , SO_4^{2-} , H_2PO_4^- , CN^- , and AcO^- . Color changes of sensor 1 and 2 are shown in Fig. 5 and 6, respectively.

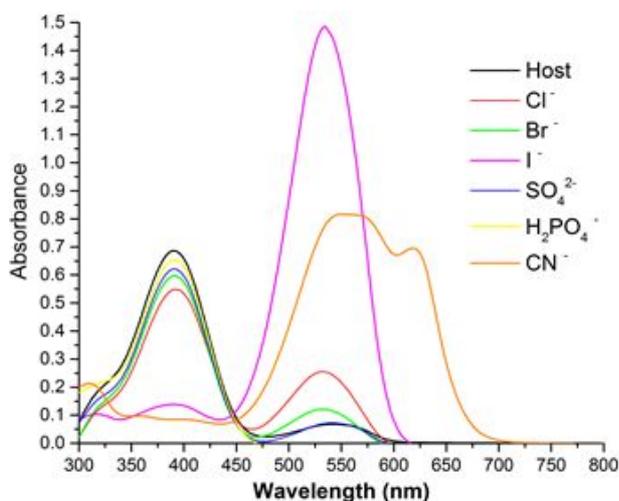


Fig 7. UV-vis spectra of 1 with various anions

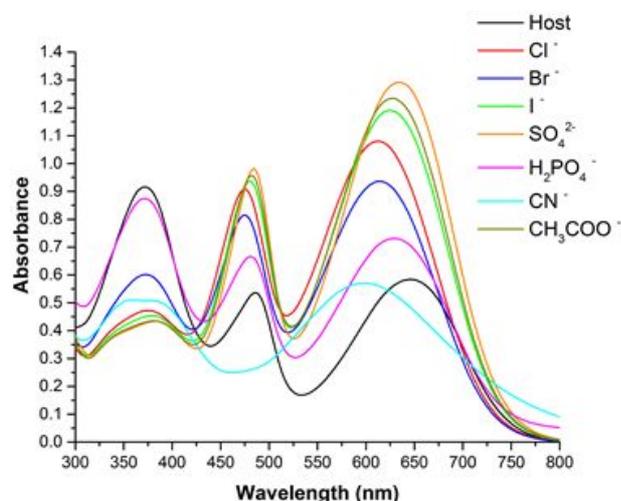


Fig 8. UV-vis spectra of 2 with various anions

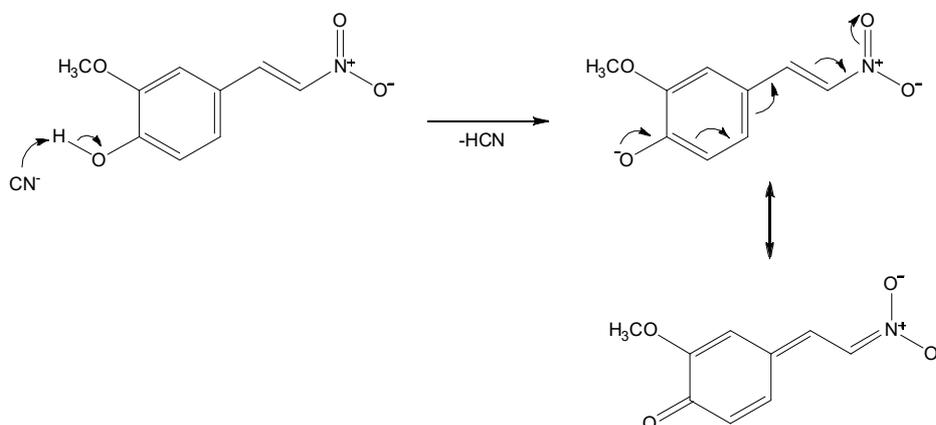


Fig 9. Prediction interaction of sensor 1 with cyanide ion

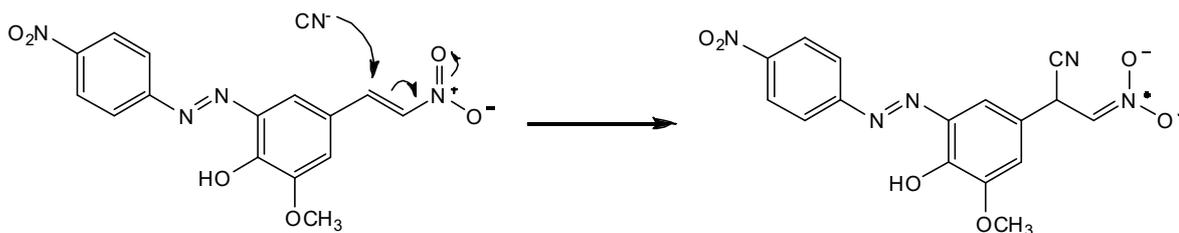


Fig 10. Prediction interaction of sensor 2 and cyanide ion

Fig 11. Color change LOD of sensor 1; a. 0 eq, b. 10^{-5} eq, c. 10^{-4} eq, d. 10^{-3} eq, e. 10^{-2} eq, f. 10^{-1} eq, g. 1 eq, h. 5 eq, i. 10 eq, j. 30 eq, k. 50 eq, l. 70 eq, m. 90 eq

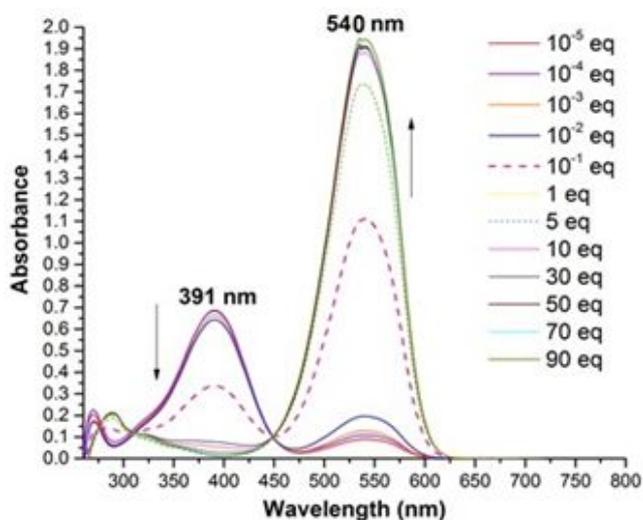


Fig 12. UV-vis spectra LOD CN^- of sensor 1

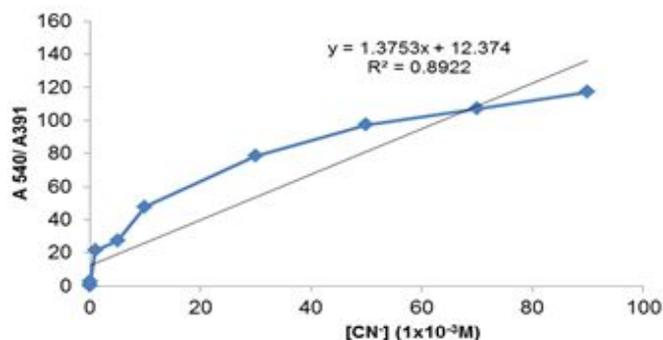


Fig 13. Job's plot between sensor 1 and cyanide

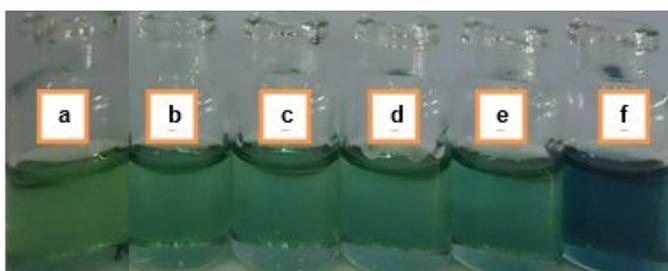


Fig 14. Color change LOD of sensor 2; a. 0 eq, b. 10^{-5} eq, c. 10^{-4} eq, d. 10^{-3} eq, e. 10^{-2} eq, f. 10^{-1} eq

The addition of three drops saturated salt anions of F^- , SO_4^{2-} , CO_3^{2-} , and AcO^- into 2 mL of **1** gave red shift (bathochromic) spectra with different intensity, whereas Cl^- , Br^- , I^- , and H_2PO_4^- did not show any significant change. Cyanide anion gave different result where the absorption band at 391 nm (ICT) completely disappeared and new absorption was shown in 620 nm. It was due to cyanide deprotonation, therefore gave bathochromic shift. Basicity of cyanide ion was the highest than others in this experiment, so it can deprotonate binding site of sensor **1** that showed

different color change. Prediction of interaction sensor **1** with cyanide ion is shown Fig. 9. Anions F^- , CO_3^{2-} , and CH_3COO^- have same change color that was red-purple, its probably caused by similar basicity. Prediction interaction that occurs between these anions and binding site of sensor **1** is hydrogen bonding. The presence of the other anions (Cl^- , Br^- , I^- , SO_4^{2-}) gave a similar color change but I^- ion showed intensity higher than others.

Limit of detection (LOD) cyanide ion in sensor **1** was measured by $3\sigma/m$ in DMSO through the UV-vis titration experiment [13-15]. Color change and UV-vis spectra are shown Fig. 11 and 12 respectively. The appearance of a distinct isosbestic point at 448 nm during the titration process indicates host-guest complexation equilibrium. The standar deviation of absorbance of sensor **1** without CN^- was obtained as $\sigma = 0.00427$ ($n = 5$). The absorbance at 391 nm and 540 nm were measured by incremental addition of CN^- to sensor **1** in various concentration, the slope of which gave $m = 1.375$. The UV-vis titration experiments of sensor **1** revealed that the LOD of CN^- was 9.316×10^{-3} M at $3\sigma/m$. Limit of detection (LOD) plot between sensor **1** and cyanide is shown in Fig. 13.

Sensor **2** gave similar color changes were seen when F^- , Cl^- , Br^- , I^- , CO_3^{2-} , SO_4^{2-} , CH_3COO^- were added that is shown in Fig. 6. In addition, cyanide and dihydrogen phosphate gave different response. Cyanide gave color change green to greenish-blue and no significant color change when dihydrogen phosphate was added. Based on UV-vis spectra that is shown in Fig. 8, sensor **2** gave similar spectral response towards H_2PO_4^- ion. The presence of the other anions (Cl^- , Br^- , and I^-) with weak basicity resulted in negligible changes in UV-vis spectra sensor **2**. Surprisingly, spectral changes were induced by addition of CN^- . Cyanide gave hypsochromic shift (blue-shift) from 647 nm to 601 nm. That phenomenon is called chemodosimeter, it is caused by nucleophilic addition CN^- to binding site of sensor **2** [15-16]. Prediction interaction of sensor **2** and CN^- is shown in Fig. 10. The binding interaction of presence others anions (F^- , Cl^- , Br^- , I^- , CO_3^{2-} , SO_4^{2-} , CH_3COO^-) is probably due to hydrogen bonding between binding site of sensor **2** and anion [17-18].

Limit of detection (LOD) color change and UV-vis spectra of sensor **2** could be seen in Fig. 14 and 15. Based on Fig. 15 there was no appearance of a distinct isosbestic point, it's mean that chemodosimeter occurred during the titration process. The standar deviation of absorbance of sensor **2** without CN^- was obtained as $\sigma = 0.10215$ ($n = 5$). The absorbance at 385 nm and 633 nm were measured by incremental addition of CN^- to sensor **2** in various concentration, the slope of which gave $m = 22.199$. The UV-vis titration

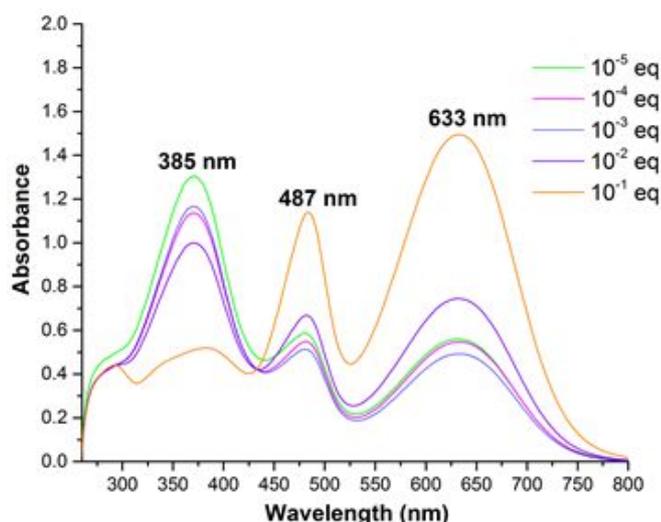


Fig 15. UV-vis spectra LOD CN^- of sensor 2

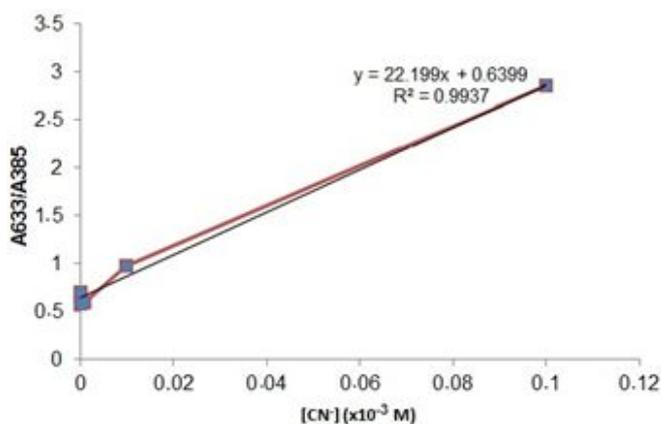


Fig 16. Job's plot between sensor 2 and cyanide

experiments of sensor 2 revealed that the LOD of CN^- was 1.370×10^{-2} M at $3\sigma/m$. Limit of detection (LOD) plot between sensor 2 and cyanide is shown in Fig. 16.

Based on the results, sensor 1 has different binding interaction with sensor 2. The color change in addition of cyanide to sensor 1 is caused by deprotonation cyanide to binding site of sensor 1. It is more favorable than nucleophilic addition in styrene group by cyanide. Therefore, significant change of color occurred from yellowish to red-purple and then purple, which could be seen in UV-vis spectra there is bathochromic shift. On the other hand, hypsochromic shift occurred in sensor 2 that was caused by nucleophilic addition (Fig. 8). Nucleophilic character of the cyanide anion enable addition of sensor 2 was favorable than deprotonation. It's probably caused by binding site of sensor 2 has steric position. There is azo group that bind nitrobenzene in ortho position of binding site, so nucleophilic addition more favorable by sensor 2 than deprotonation.

CONCLUSION

In summary, we have successfully synthesized a colorimetric anion sensor from vanillin. Sensor 1 and 2 gave selectivity for cyanide, followed by color changing from yellowish to red-purple (sensor 1) and green to greenish-blue (sensor 2). Color change that occur in sensor 1 is due to the deprotonation binding site, whereas in sensor 2 is caused by nucleophilic addition of cyanide.

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